EQUILIBRIUM CLASSES

Sample Paper – 2014
Class – XI
Subject – Chemistry

TOPIC – Atomic Structure & Same Basic Concept Of Chemistry

Attempt all questions.

Q-1 (a) Write electronic Configuration of Cr. & Cu in s.p.d.f.
(b) What is the value of m, s, and l for n=3   [4]

Q-2 (a) What is the de Broglie's equation and explain dual nature of electron.
(b) Calculate the Value of n, l, m & s for last electron of Zn and Vanadium.   [4]

Q-3 Write Short notes on-
   (a) Heisenberg uncertainty principle.
   (b) A Chemical Compound has following% composition-
   K = 22.8%, Fe = 15.2%, C = 19.5%, N = 22.8%
   Calculate simple empirical formula.   [4]

Q-4 (a) 2.8g of KOH is dissolved in water to give 200cc of solution. Calculate the molarity of KOH in solution.   [2]
(b) Calculate the molarity of 20% solution of KOH by mass where density is $1.02ml^{-1}$.   [2]

Q-5 (a) How many electrons are present in 1 kg.   [1]
(b) What will be volume of H_2 measured at STP that would be liberated when 8 g of calcium is completely reacts with water.

$$Ca + 2H_2O \rightarrow Ca(OH)_2 + H_2$$ [3]

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Q-6 Why 2d orbital is not possible.

Q-7 What is Hydrogen Spectrum explain with Lyman, Balmer, Pauchen & Pfund series.

Q-8 Explain Azimuthal Quantum Number.

Q-9 Write main postulates of Bohr's Model.

Q-10 Explain the limitations of Rutherford Model.

Our aim is to give a right path towards success
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