## Guess Paper – 2014
## Class – XII
## Subject –Chemistry

| Q1. | The standard enthalpy of formation. | 2 |
|---|---|---|
| Q2. | Density of a gas is found to be 5.46 g/dm$^3$ at 27 °C at 2 bar pressure. What will be its density at STP? | 2 |
| Q3. | Why do gases deviate from the ideal behaviour?Under what condition real gases behave ideally?<br> Or<br>Define Dalton's Law of partial pressure. | 2 |
| Q4. | Define dipole-induce dipole force of attraction with one example. | 2 |
| Q5. | Define the First Law of thermodynamics. | 2 |
| Q6. | For the reaction at 298 K,<br>$2A_{(g)} + B_{(g)} \rightarrow C_{(g)}$<br>$\Delta°H = 400$ kJ mol$_{-1}$ and $\Delta°S = 0.2$ kJ K$^{-1}$mol$^{-1}$<br>At what temperature will the reaction become spontaneous considering $\Delta°H$ and $\Delta°S$ to be constant over the temperature range. | 2 |
| Q7. | What are hybridisation states of each carbon atom calculate sigma and pi bond in given compound? $(CN)_3CCH=CHCOCH_3$ | 2 |
| Q8. | I<br>)Give the IUPAC names of the following compound.<br>$CH_3CH=CHCOOH$<br>ii) Draw the structure of Pent-4-en-2-ol. | 2 |
| Q9. | Draw the resonance structures for the following compounds. Show the electron shift using curved-arrow notation.<br>(a) $C_6H_5OH$ (b) $C_6H_5NO_2$<br>(b) | 2 |
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| Q10. | Write the difference between following with one example in each:<br>i)Homolytic bond fission and heterolytic bond fission.<br>ii)Aromatic and anti aromatic compound. | 3 |
|---|---|---|
| Q11. | i)Define $P_C$ , $V_C$ and $T_C$ .<br>ii)Define copresibility factor. | 4 |

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