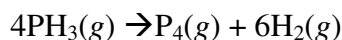


CBSE Class 12th Chemistry

General instructions:

- 1) All questions are compulsory.
- 2) Marks for each question are indicated against it.
- 3) Question number 1 to 8 are very short –answer questions, carrying 1 mark each. Answer these in one word or about one sentence each.
- 4) Question number 9 to 18 are short –answer questions, carrying 2 marks each. Answer these in about 30 words each.
- 5) Question number 19 to 27 are short –answer questions, carrying 3 marks each. Answer these in about 40 words each.
- 6) Question number 28 to 30 are long-answer questions of 5 marks each. Answer these in about 70 words each.
- 7) Use log tables, if necessary. Use of calculators is not permitted.

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|------|--|---|
| Q1. | Give the IUPAC name of the following compound.
(CH ₃) ₃ CCOC(CH ₃) ₂ CHBr CH ₃ | 1 |
| Q2 | Draw the structures of the 1-Bromo-4-sec. butyl-2-methylbenzene . | 1 |
| Q3. | What happens when Gluconic acid is treated with nitric acid? | 1 |
| Q4. | Define Zeta potential. | 1 |
| Q5 | Define synergic bonding. | 1 |
| Q6. | Write the reaction of XeF ₂ with PF ₅ . | 1 |
| Q7. | Write reaction between 1-methoxy propene with HI followed by dilute NaOH. | 1 |
| Q8. | Define Rate constant . | 1 |
| Q9. | Why molar conductivity increases and conductivity decreases sharply with increase in dilution for weak electrolyte? | 2 |
| Q10. | Explain the following terms with suitable examples:
(i) Why amorphous solid are isotropic and crystalline solid are anisotropic in nature .
(ii) What happens when Fe ₃ O ₄ heated above 850K? | 2 |
| Q11. | An element with molar mass 2.7×10 ⁻² kg mol ⁻¹ forms a cubic unit cell with edge length 405 pm. If its density is 2.7×10 ³ kg ⁻³ , Calculate the number of atoms and unit cells in 150 g of element? | 2 |
| Q12. | a) What is the difference between molecularity and order of reaction?
b) Define temperature coefficient of a reaction. | 2 |
| Q13. | Complete the following reaction:
i) XeF ₂ + H ₂ O →
ii) Pt + HNO ₃ (conc) + HCl → | 2 |
| Q14. | The decomposition of phosphine, PH ₃ , proceeds according to the following equation: | 2 |



It is found that the reaction follows the following rate equation:

Rate = $k[\text{PH}_3]$ if initial pressure was 0.5 bar and total pressure at 200 sec was 1.5 bar at 100°C.

The half-life of PH_3 is 37.9 s at 120°C. calculate E_a .

Q15. i) Draw the structure of Cyclotrimetaphosphoric acid. 2
 ii) Why Perchloric acid is stronger than sulphuric acid?

Q16. Draw all the isomers (geometrical and optical) of: (i) $[\text{CoCl}_2(\text{en})_2]^+$, (ii) $[\text{Co}(\text{NH}_3)\text{Cl}(\text{en})_2]^{2+}$ (iii) $[\text{Co}(\text{NH}_3)_2\text{Cl}_2(\text{en})]^+$ 2

Or

Giving a suitable example for each, explain the following:

(i) Tetrahedral Crystal field splitting of d subshell

(ii) Linkage isomerism

(iii) Ambidentate ligand.

Q17. i) NH_2 group of aniline acetylated is before carrying out nitration? 2

ii) Aniline does not undergo Friedel-Crafts reaction.

Q18. i) Write a short note on Hofmann's bromamide reaction. 2

ii) Write a short note on Carbyl amine reaction.

Q19. i) Write mechanism of hydration of ethene. 3

ii) Convert methanol to butan-1-ol.

Q20. A cell, Ag/Ag^+ (saturated Ag_2CrO_4 solution) \parallel Ag^+ (0.1 molar) / Ag , if EMF is 0.164 V, then calculate the following: (Given $E^\circ \text{Ag}^+/\text{Ag} = 0.80 \text{ V}$) 3

i) Concentration of Ag_2CrO_4 solution

ii) Solubility product of Ag_2CrO_4 solution

iii) E° cell

Or

Three electrolytic cells A, B, C containing solutions of ZnSO_4 , AgNO_3 and CuSO_4 , respectively are connected in series. A steady current of 1.5 amperes was passed through them until 1.45 g of silver deposited at the cathode of cell B. How long did the current flow? What mass of copper and zinc were deposited?

(At. mass $\text{Cu} = 63.5, \text{Ag} = 108, \text{Zn} = 65.39$)

Q21. i) What happens when a colloidal sol of AgI is mixed with that of AgNO_3 sol? 3

ii) What is the difference between gel and sol?

iii) Write the cause of peptization.

Q22. Describe the role of the following: 3

(i) NaCN in the extraction of Ag .

(ii) I_2 in the purification of Zirconium.

(iii) Leaching in the extraction of aluminium?

Q23. i) Why fluorine does not show +ve oxidation state? 3

ii) What happens when PCl_5 react with heavy water?

iii) How can you detect presence of NO_3^- ?

Q24. i) Define enantiomer. Calculate no. of optically active stereo isomers in 2,3 dichloro butane. 3

- ii) Arrange the following in increasing order of reactivity towards S_N^1 :
 $(CH_3)_2CHI, CH_2=CHCH_2I, C_6H_5CH_2I$
- iii) Prepare 1-Fluoro propane from 2-Chloropropane
- Q25. i) What is tincture of iodine? What is its use? 3
 ii) Write the composition of Dettol soap.
 iii) Write uses of eucalypt.
- Q26. Pradeep had very high fever. He was given strong antibiotics. But after recovering from fever he was not able to digest food and was feeling too weak. The grandmother who lived in his neighborhood suggested him to take lots of milk, egg, fruits and vegetables. 3
 a. Why?
 b. What is the remedy for this?
 c. What was the value that Pradeep had by taking fruits and vegetables?
- Q27. i) Distinguish between the terms thermoplastic and thermosetting plastic and give an example of each. 3
 ii) Write the monomer of nylon 6.
 iii) Write the monomer of glyptal.
- Q28. i) The actinoids exhibit more number of oxidation states in general than the lanthanoids. 5
 ii) Write the cause and consequences of lanthanoid contraction.
 iii) With same (d^4) configuration Cr (II) is reducing whereas Mn (III) is oxidising.
 iv) How to prepare $K_2Cr_2O_7$ from chromite ore.
 v) Why trend of melting point is minimum and oxidation number is maximum in middle of transition series.
- Or
- (a) Complete the following chemical equations:
 (i) $Fe^{3+} + I^- + H^+ \rightarrow$
 (ii) $Cr_2O_7^{2-} + Sn^{2+} + H^+ \rightarrow$
- (i) Why is Eu (III) more stable than Ce (III)?
 iv) On the basis of data given below comment on the stability of Fe^{3+} in acid solution as compared to that of Cr^{3+} or Mn^{3+} .
 $[Cr^{2+}/Cr = -0.9V, Cr^{3+}/Cr^{2+} = -0.4V, Mn^{2+}/Mn = -1.2V, Mn^{3+}/Mn^{2+} = +1.5V, Fe^{2+}/Fe = -0.4V, Fe^{3+}/Fe^{2+} = +0.8V]$
- v) Write the reaction of Mn²⁺ ion with peroxydisulphuric acid.
- Q29. i) Write the mechanism of esterification. 5
 ii) Distinguish between methyl benzoate and isopropyl benzoate.
 iii) An organic compound (A) with molecular formula C_6H_{12} when treated with O_3 , zinc and water gives compound B & C both forms an orange-red precipitate with 2,4-DNP reagent and gives yellow precipitate on heating with iodine in the presence of sodium hydroxide, B reduces Tollens' or Fehlings' reagent but C does not reduce Tollens' or Fehlings' reagent, B and C on treating with diluted alkali followed by heating gives D and E are isomers of each other decolourise bromine water or Baeyer's reagent, but only D ($C_6H_{10}O$) can give iodoform test to give (F) $[C_5H_7O_2Na]$. Identify the compounds (A), (B), (C), D, E and explain the reactions involved.

Or

- i) Arrange the following compounds in increasing order of their reactivity in nucleophilic addition reactions: Acetaldehyde, Acetone, Benzaldehyde, *p*-Tolualdehyde, *p*-Nitrobenzaldehyde, Acetophenone.
- ii) Convert Benzaldehyde to cinnamic acid (3-phenyl prop-2 enoic acid) in two steps.
- iii) Describe Cannizzaro reaction with example.
- iv) Although phenoxide ion has more number of resonating structures than carboxylate ion, carboxylic acid is a stronger acid than phenol. Why?
- v) Ethanal to 3-hydroxy butanoic acid in two steps.
- Q30. a) Define cryoscopic constant, write its unit.
- b) Define Henry law. Write one application.
- c) 72.5 g of phenol is dissolved in 1 kg of a solvent ($k_f = 14$) which leads to dimerization of phenol and freezing point is lowered by 7 kelvin. What percent of total phenol is present in dimeric form?
- Or
- a) What is meant by negative deviation from Raoult's law? Draw diagram to illustrate the relationship between vapour pressure and mole fraction of carbon disulphide and acetone in a solution to represent negative or positive deviation?
- b) 1.22 g of benzoic acid is dissolved in (i) 100 g acetone (K_b for acetone = 1.7) and (ii) 100 g benzene (K_b for benzene = 2.6). The elevation in boiling points T_b is 0.17°C and 0.13°C respectively.
- (a) What are the molecular weights of benzoic acid in both the solutions?
- (b) What do you deduce out of it in terms of structure of benzoic acid?