
General Instructions:

1. All Questions are compulsory.
 2. Marks for each question are indicated against it.
 3. Question number 1 to 8 are very short answer questions and carry 1 mark each.
 4. Question number 9 to 18 are short answer questions and carry 2 marks each
 5. Question number 19 to 27 are also short answer questions and carry 3 marks each.
 6. Question number 28 to 30 are long answer questions and carry 5 marks each.
 7. There will be no overall option. Internal choice is given for all three 5 marks questions.
 8. Use log table if necessary, use of calculator is not permitted.
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Q1 Define Pseudo first order reaction with example?

Q2 What do you mean by state selective catalysis?

Q3 Draw the shape of BrF_3 ?

Q4 Write the IUPAC name of the compound $[\text{Co}(\text{NH}_3)_3(\text{ONO})]^{2+}$

Q5 Write the formula and chemical name of DDT.

Q6 Convert propene to acetone?

Q7 Explain HVZ reaction?

Q8 Distinguish between Propanol and Propanone?

Q9 How would you account the following

(1) Frankel defect is not found in alkali metal halides.

(2) Schotkey defects lower density of the solid.

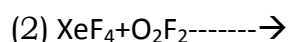
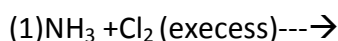
Q10 Analysis show that nickel oxide has formula $\text{Ni}_{0.98}\text{O}_{1.00}$. What fraction of Nickel exists as Ni^{2+} and Ni^{3+} .

Q11 a) The molar conductivity of acetic acid solution at infinite dilution is $390.7 \Omega^{-1}\text{cm}^2\text{mol}^{-1}$. Calculate the molar conductivity of 0.01M acetic acid solution given that the dissociation constant of acetic acid is 1.8×10^{-5}

Q12 Write difference between molecularity and order of reaction?

Q13 Derive the relationship between half life of first order reaction and its rate constant?

Q14. Complete the following:



Q15 Draw the structure of following:

1. XeOF_4 2. H_3PO_3

Q16 Explain on the basis of Valence bond theory that $[\text{Ni}(\text{CN})_4]^{2-}$ ion with square planar structure is diamagnetic and the $[\text{NiCl}_4]^{2-}$ ion with tetrahedral geometry is paramagnetic

Q17 Arrange the following order of properties mentioned

$\text{C}_2\text{H}_5\text{NH}_2$, $\text{C}_6\text{H}_5\text{NH}_2$, NH_3 , $(\text{C}_2\text{H}_5)_2\text{NH}$ (Basic Strength)

$\text{C}_2\text{H}_5\text{NH}_2$, $(\text{C}_2\text{H}_5)_3\text{N}$, NH_3 , $(\text{C}_2\text{H}_5)_2\text{NH}$ (Basic Strength in gases phase)

Q18 (1) Give one test to distinguish between Methylamine and dimethylamine?

(2) Primary amine have high boiling point than comparable tertiary amine?

Q19 (1) What happened when Dglucose is treated with (1).HI, (2) HNO_3 2

(2) What is the difference between a nucleotide and nucleoside? 1

OR

1. What do means by essential and non essential amino acid Give an example? 2

2. Explain Zwitter ion structure of amino acid? 1

Q20(1) Write the mechanism of free radical for the polymerization of ethane? 2

(2) Write monomer of following

(1).Teflon (2) bakelite (3) PVC (4) N-66

Q21 Discuss with example biodegradable and non biodegradable detergents with example?
3

OR

Explain:

(a) Why is the use of aspartame limited to cold food and drinks?

(b)What problem arises in using Alitame as artificial sweetener?

(C) Explain anti fertility drugs? 3

Q22 (1)Write the name of the electrolyte used in (i) fuel cell (ii) mercury cell.

(2)Predict, if the reaction between $\text{Fe}^{3+}(\text{aq})$ and $\text{I}^{-}(\text{aq})$. It is given that

$$E^{\circ}_{\text{Fe}^{3+}/\text{Fe}^{2+}} = +0.77\text{v} \quad E^{\circ}_{\text{I}_2/\text{I}^{-}} = 0.54\text{v}$$

Q23(1)What is adsorption isotherm .Describe Freundlich adsorption isotherm?

(2)Which of the following electrolyte is most effective for the coagulation of $\text{Fe}(\text{OH})_3$ sol and why NaCl , Na_2SO_4 Na_3PO_4 .

Q24Describe the principle of the following:

Van Arkel method, Zone process, electrolytic refining

Q25. 1. Why does PCl_3 fumes in the air?

2 . What is the basicity of H_3PO_4 ?

3. How is O_3 estimated quantitatively

OR

Prepare the nitric acid and sulphuric acid by Ostwald and CONTACT PROCESS ?

Q26 (a) Explain why

(1) Vinyl chloride is unreactive in nucleophilic substitution reactions.

(2) The dipole moment of chlorobenzene is lower than cyclohexyl chloride

(3)What happened when propene is treated with HBr in the presence of Peroxide?

Q27(1) Write chemical reaction for Riemer –Tieman reaction, Sandmeyer reaction

(2) Give distinction between primary, secondary, tertiary alcohol by Victor Meyer test?

Or

(1) Convert Ethane to methanol

(2) Write the mechanism of hydration of ethane to yield ethanol

Q28 (1) 2g benzoic acid dissolved in 25g of benzene show a depression in freezing point equal to 1.62K. Molar depression constant for benzene is 4.9Kkgmol^{-1} . What is the percentage associated of acid if it forms double molecule in solution?

(2) Define the following: 1 Molarity 2 Molality and which one is better and why?

OR (1) Two elements A and B form compounds having molecular formula AB_2 and AB_4 . When dissolved in 20g of benzene, 1g of AB_2 lowers the freezing point by 2.3K, whereas 1.0g of AB_4 lowers it by 1.3K. The molar depression constant for benzene is 5.1Kkgmol^{-1} . Calculate atomic mass of A and B?

(2) Define Henry's Law and explain its two applications?

Q29. (1) What is the effect of pH on the color of the solution of potassium dichromate?

2 Why are the compounds of transition elements coloured?

3 Why do transition elements act as catalysts? Give two examples.

4. Why do transition elements form (a) interstitial compounds and (b) Alloys?

5 Why are Ni^{2+} compounds thermodynamically more stable than Pt^{2+} compounds, whilst Pt^{4+} compounds are relatively more stable than Ni^{4+} compounds?

OR

Name a transition metal which does not exhibit variation in oxidation state in its compounds.

or

Assign reason for each of the following:

(1) Ce^{3+} can be easily oxidised to Ce^{4+} .

(2) E° for $\text{Mn}^{3+}/\text{Mn}^{2+}$ couple is more positive than for $\text{Fe}^{3+}/\text{Fe}^{2+}$.

(3) Transition metals exhibit higher enthalpies of atomization.

(4) Differentiate the properties of Lanthanoids and actinoids?

(5) Describe with chemical reaction for the preparation of potassium permanganate from pyrolusite ore.

Q30 (1) An organic compound (A) with molecular formula C_8H_8O forms an orange-red precipitate with 2,4-DNP reagent and gives a yellow precipitate on heating with iodine in the presence of sodium hydroxide. It neither reduces Tollens' reagent or Fehling's reagent, nor does it decolorize bromine water or Bayer's reagent. On drastic oxidation with a chromic acid it gives a carboxylic acid (B) having molecular formula $C_7H_6O_2$. Identify (A) and (B) and explain the reaction involved?

(2) (ii) Cannizzaro's reaction

(iii) Aldol condensation

(b) Of aldehydes and ketones which would reduce Tollens' reagent? Why

OR

(I) How will you distinguish between the following pairs of compounds by chemical tests?

i) Propanal and Propanone

ii) Acetophenone and benzophenone

iii) Phenol and benzoic acid

(II) How will you convert in not more than two steps-

(i) Propanone to propene

(ii) benzoic acid to benzaldehyde

ALL THE BEST

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