



TEST: 4 CHEMISTRY –XII CHAPTER :-p-block Group-17 elements

Q.1 Draw the structure for  $\text{HClO}_3$ .

Q.2 Draw the structure of:  $\text{BrF}_5$

OR

Draw the structure of following:  $\text{ClF}_3$ .

Q.3 Arrange the following into decreasing order of their reducing character  $\text{HF}$ ,  $\text{HCl}$ ,  $\text{HBr}$ ,  $\text{HI}$ .

Q.4 Account for the following :

Fluorine forms only one oxoacid  $\text{HOF}$ .

Q.5 Account for the following:

Acidic character increases from  $\text{HF}$  to  $\text{HI}$ .

Q.6  $\text{F}_2$  has lower bond dissociation enthalpy than  $\text{Cl}_2$ . Why?

Q.7 Account for the following :

Interhalogens are more reactive than pure halogens.

Q.8 Why are pentahalides of a metal more covalent than its trihalides?

Q.9  $\text{ICl}$  is more reactive than  $\text{I}_2$ .

Q.10 Draw the structure of  $\text{HClO}_4$ .

Q.11 Draw the structure of  $\text{HOClO}_2$  molecule.

Q.12 Complete the following chemical equation.



Q.13 Complete the following chemical equation.



Q.14 Why does fluorine not play the role of a central atom in interhalogen compounds?

Q.15 How are interhalogen compounds formulated and how are they prepared?

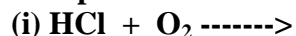
Or

How are interhalogen compounds formed? What general compositions can be assigned to them?

Q.16 Draw the structure of:

(i) Hypochlorous acid (ii) Chlorous acid

Q.17 Complete the reaction:



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